

# Chemistry Data Booklet

## National 5

For use in National Qualification Courses

Publication date: 2021

Publication code: BB8395

Published by the Scottish Qualifications Authority  
The Optima Building, 58 Robertson Street, Glasgow G2 8DQ  
Lowden, 24 Wester Shawfair, Dalkeith, Midlothian EH22 1FD

[www.sqa.org.uk](http://www.sqa.org.uk)

*The information in this publication may be reproduced in support of SQA qualifications. If it is reproduced, SQA should be clearly acknowledged as the source. If it is to be used for any other purpose, then written permission must be obtained from SQA. It must not be reproduced for trade or commercial purposes.*

For an up-to-date list of prices visit the **Publication Sales and Downloads** section of SQA's website.  
For further details telephone SQA's Customer Contact Centre on 0345 279 1000.

## Contents

Relationships for National 5 Chemistry and Specific Heat Capacity of Water . . . . .	3
Periodic Table of the Elements. . . . .	4
Melting and Boiling Points of Selected Elements. . . . .	5
Electron Arrangements of Main Group Elements. . . . .	6
Flame Colours . . . . .	6
Names, Symbols, Relative Atomic Masses, Densities and Dates of Discovery . . . . .	7
Formulae of Selected Ions containing more than one kind of Atom. . . . .	8
Solubilities of Selected Compounds in Water . . . . .	8
Melting and Boiling Points of Selected Inorganic Compounds . . . . .	9
Melting and Boiling Points of Selected Organic Compounds . . . . .	9
Electrochemical Series (Reduction Reactions). . . . .	10

### Relationships for National 5 Chemistry

$$E_h = cm\Delta T$$

$$n = cV$$

$$n = \frac{m}{GFM}$$

$$\frac{c_1V_1}{n_1} = \frac{c_2V_2}{n_2}$$

$$\text{average rate} = \frac{\Delta \text{quantity}}{\Delta t}$$

$$\% \text{ by mass} = \frac{m}{GFM} \times 100$$

### Specific Heat Capacity of Liquid Water

$$c = 4.18 \text{ kJ kg}^{-1} \text{ }^\circ\text{C}^{-1}$$

# Periodic Table of the Elements

Column 1	Column 2											Column 3	Column 4	Column 5	Column 6	Column 7	Column 0										
1 Hydrogen <b>H</b>		<div style="display: flex; align-items: center; justify-content: center;"> <div style="margin-right: 20px;"><b>Key</b></div> <div style="border: 1px solid black; padding: 5px;">                 Atomic Number Name of Element Symbol             </div> </div>																									2 Helium <b>He</b>
3 Lithium <b>Li</b>	4 Beryllium <b>Be</b>																										
11 Sodium <b>Na</b>	12 Magnesium <b>Mg</b>	<b>TRANSITION METALS</b>										13 Aluminium <b>Al</b>	14 Silicon <b>Si</b>	15 Phosphorus <b>P</b>	16 Sulfur <b>S</b>	17 Chlorine <b>Cl</b>	18 Argon <b>Ar</b>										
19 Potassium <b>K</b>	20 Calcium <b>Ca</b>	21 Scandium <b>Sc</b>	22 Titanium <b>Ti</b>	23 Vanadium <b>V</b>	24 Chromium <b>Cr</b>	25 Manganese <b>Mn</b>	26 Iron <b>Fe</b>	27 Cobalt <b>Co</b>	28 Nickel <b>Ni</b>	29 Copper <b>Cu</b>	30 Zinc <b>Zn</b>	31 Gallium <b>Ga</b>	32 Germanium <b>Ge</b>	33 Arsenic <b>As</b>	34 Selenium <b>Se</b>	35 Bromine <b>Br</b>	36 Krypton <b>Kr</b>										
37 Rubidium <b>Rb</b>	38 Strontium <b>Sr</b>	39 Yttrium <b>Y</b>	40 Zirconium <b>Zr</b>	41 Niobium <b>Nb</b>	42 Molybdenum <b>Mo</b>	43 Technetium <b>Tc</b>	44 Ruthenium <b>Ru</b>	45 Rhodium <b>Rh</b>	46 Palladium <b>Pd</b>	47 Silver <b>Ag</b>	48 Cadmium <b>Cd</b>	49 Indium <b>In</b>	50 Tin <b>Sn</b>	51 Antimony <b>Sb</b>	52 Tellurium <b>Te</b>	53 Iodine <b>I</b>	54 Xenon <b>Xe</b>										
55 Caesium <b>Cs</b>	56 Barium <b>Ba</b>	57 Lanthanum <b>La</b>	58–71 ●	72 Hafnium <b>Hf</b>	73 Tantalum <b>Ta</b>	74 Tungsten <b>W</b>	75 Rhenium <b>Re</b>	76 Osmium <b>Os</b>	77 Iridium <b>Ir</b>	78 Platinum <b>Pt</b>	79 Gold <b>Au</b>	80 Mercury <b>Hg</b>	81 Thallium <b>Tl</b>	82 Lead <b>Pb</b>	83 Bismuth <b>Bi</b>	84 Polonium <b>Po</b>	85 Astatine <b>At</b>	86 Radon <b>Rn</b>									
87 Francium <b>Fr</b>	88 Radium <b>Ra</b>	89 Actinium <b>Ac</b>	90–103 ■	104 Rutherfordium <b>Rf</b>	105 Dubnium <b>Db</b>	106 Seaborgium <b>Sg</b>	107 Bohrium <b>Bh</b>	108 Hassium <b>Hs</b>	109 Meitnerium <b>Mt</b>	110 Darmstadtium <b>Ds</b>	111 Roentgenium <b>Rg</b>	112 Copernicium <b>Cn</b>	113 Nihonium <b>Nh</b>	114 Flerovium <b>Fl</b>	115 Moscovium <b>Mc</b>	116 Livermorium <b>Lv</b>	117 Tennessine <b>Ts</b>	118 Oganesson <b>Og</b>									

	58 Cerium <b>Ce</b>	59 Praseodymium <b>Pr</b>	60 Neodymium <b>Nd</b>	61 Promethium <b>Pm</b>	62 Samarium <b>Sm</b>	63 Europium <b>Eu</b>	64 Gadolinium <b>Gd</b>	65 Terbium <b>Tb</b>	66 Dysprosium <b>Dy</b>	67 Holmium <b>Ho</b>	68 Erbium <b>Er</b>	69 Thulium <b>Tm</b>	70 Ytterbium <b>Yb</b>	71 Lutetium <b>Lu</b>
	90 Thorium <b>Th</b>	91 Protactinium <b>Pa</b>	92 Uranium <b>U</b>	93 Neptunium <b>Np</b>	94 Plutonium <b>Pu</b>	95 Americium <b>Am</b>	96 Curium <b>Cm</b>	97 Berkelium <b>Bk</b>	98 Californium <b>Cf</b>	99 Einsteinium <b>Es</b>	100 Fermium <b>Fm</b>	101 Mendelevium <b>Md</b>	102 Nobelium <b>No</b>	103 Lawrencium <b>Lr</b>

Elements below the dark line are metals.

## Melting and Boiling Points of Selected Elements

Group 1	Group 2											Group 3	Group 4	Group 5	Group 6	Group 7	Group 0												
1 Hydrogen -259 -253		<div style="display: flex; align-items: center; justify-content: center;"> <div style="margin-right: 10px;">Key</div> <div style="border: 1px solid black; padding: 5px;">                     Atomic Number                      Name of Element                      Melting Point (°C)  <i>Boiling Point (°C)</i> </div> </div>																											2 Helium -272 -269
3 Lithium 181 1342	4 Beryllium 1287 2471*																					5 Boron 2077 4000	6 Carbon †3825	7 Nitrogen -210 -196	8 Oxygen -219 -183	9 Fluorine -220 -188	10 Neon -249 -246		
11 Sodium 98 883	12 Magnesium 650 1090																					13 Aluminium 660 2519	14 Silicon 1414 3265	15 Phosphorus 44 281	16 Sulfur 115 445	17 Chlorine -101 -34	18 Argon -189 -186		
19 Potassium 63 759	20 Calcium 842 1484	21 Scandium 1541 2836	22 Titanium 1670 3287	23 Vanadium 1910 3407	24 Chromium 1907 2671	25 Manganese 1246 2061	26 Iron 1538 2861	27 Cobalt 1495 2927	28 Nickel 1455 2913	29 Copper 1085 2560	30 Zinc 420 907	31 Gallium 30 2229	32 Germanium 938 2833	33 Arsenic *817 †616	34 Selenium 221 685	35 Bromine -7 59	36 Krypton -157 -153												
37 Rubidium 39 688	38 Strontium 777 1377	39 Yttrium 1522 3345	40 Zirconium 1854 4406	41 Niobium 2477 4741	42 Molybdenum 2622 4639	43 Technetium 2157 4262	44 Ruthenium 2333 4147	45 Rhodium 1963 3695	46 Palladium 1555 2963	47 Silver 962 2162	48 Cadmium 321 767	49 Indium 157 2072	50 Tin 232 2586	51 Antimony 631 1587	52 Tellurium 450 988	53 Iodine 114 184	54 Xenon -112 -108												
55 Caesium 28 671	56 Barium 727 1845	57 Lanthanum 920 3464	72 Hafnium 2233 4600	73 Tantalum 3017 5455	74 Tungsten 3414 5555	75 Rhenium 3185 5590	76 Osmium 3033 5008	77 Iridium 2446 4428	78 Platinum 1768 3825	79 Gold 1064 2836	80 Mercury -39 357	81 Thallium 304 1473	82 Lead 327 1749	83 Bismuth 271 1564	84 Polonium 254 962	85 Astatine 302	86 Radon -71 -62												

\* at 28 atmospheres  
 † sublimates

## Electron Arrangements of Main Group Elements

Key

Atomic Number Name of Element Symbol Electron arrangement
--------------------------------------------------------------------

Group 1	Group 2	Group 3	Group 4	Group 5	Group 6	Group 7	Group 0
1 Hydrogen H 1							2 Helium He 2
3 Lithium Li 2,1	4 Beryllium Be 2,2	5 Boron B 2,3	6 Carbon C 2,4	7 Nitrogen N 2,5	8 Oxygen O 2,6	9 Fluorine F 2,7	10 Neon Ne 2,8
11 Sodium Na 2,8,1	12 Magnesium Mg 2,8,2	13 Aluminium Al 2,8,3	14 Silicon Si 2,8,4	15 Phosphorus P 2,8,5	16 Sulfur S 2,8,6	17 Chlorine Cl 2,8,7	18 Argon Ar 2,8,8
19 Potassium K 2,8,8,1	20 Calcium Ca 2,8,8,2	31 Gallium Ga 2,8,18,3	32 Germanium Ge 2,8,18,4	33 Arsenic As 2,8,18,5	34 Selenium Se 2,8,18,6	35 Bromine Br 2,8,18,7	36 Krypton Kr 2,8,18,8
37 Rubidium Rb 2,8,18, 8,1	38 Strontium Sr 2,8,18, 8,2	49 Indium In 2,8,18, 18,3	50 Tin Sn 2,8,18, 18,4	51 Antimony Sb 2,8,18, 18,5	52 Tellurium Te 2,8,18, 18,6	53 Iodine I 2,8,18, 18,7	54 Xenon Xe 2,8,18, 18,8
55 Caesium Cs 2,8,18, 18,8,1	56 Barium Ba 2,8,18, 18,8,2	81 Thallium Tl 2,8,18, 32,18,3	82 Lead Pb 2,8,18, 32,18,4	83 Bismuth Bi 2,8,18, 32,18,5	84 Polonium Po 2,8,18, 32,18,6	85 Astatine At 2,8,18, 32,18,7	86 Radon Rn 2,8,18, 32,18,8
87 Francium Fr 2,8,18,32, 18,8,1	88 Radium Ra 2,8,18,32, 18,8,2						

← The elements on this side of the dark line are metals.
The elements on this side of the dark line are non-metals. →

### Flame Colours

Element	Ion	Flame colour
barium	Ba <sup>2+</sup>	green
calcium	Ca <sup>2+</sup>	orange-red
copper	Cu <sup>2+</sup>	blue-green
lithium	Li <sup>+</sup>	red

Element	Ion	Flame colour
potassium	K <sup>+</sup>	lilac
sodium	Na <sup>+</sup>	yellow
strontium	Sr <sup>2+</sup>	red

## Names, Symbols, Relative Atomic Masses, Densities and Dates of Discovery

(Relative atomic masses, also known as average atomic masses, have been rounded to the nearest 0.5)

Element	Symbol	Relative atomic mass	Density (g cm <sup>-3</sup> )	Date of Discovery
Actinium	Ac	227	10.1	1899
Aluminium	Al	27	2.70	1825
Americium	Am	243	12.0	1944
Antimony	Sb	122	6.68	Ancient
Argon	Ar	40	0.0018	1894
Arsenic	As	75	5.75	~1250
Astatine	At	210	unknown	1940
Barium	Ba	137.5	3.62	1808
Berkelium	Bk	247	13.3	1949
Beryllium	Be	9	1.85	1798
Bismuth	Bi	209	9.79	1753
Boron	B	11	2.34	1808
Bromine	Br	80	3.10	1826
Cadmium	Cd	112.5	8.69	1817
Calcium	Ca	40	1.54	1808
Californium	Cf	251	15.1	1950
Carbon	C	12	*	Prehistoric
Cerium	Ce	140	6.77	1803
Caesium	Cs	133	1.87	1860
Chlorine	Cl	35.5	0.0032	1774
Chromium	Cr	52	7.15	1797
Cobalt	Co	59	8.86	1739
Copper	Cu	63.5	8.96	Ancient
Curium	Cm	247	13.5	1944
Dysprosium	Dy	162.5	8.55	1886
Einsteinium	Es	252	unknown	1952
Erbium	Er	167.5	9.07	1843
Europium	Eu	152	5.24	1896
Fluorine	F	19	0.0017	1886
Francium	Fr	223	unknown	1939
Gadolinium	Gd	157	7.90	1880
Gallium	Ga	69.5	5.91	1875
Germanium	Ge	72.5	5.32	1886
Gold	Au	197	19.3	Ancient
Hafnium	Hf	178.5	13.3	1923
Helium	He	4	0.0002	1868
Holmium	Ho	165	8.80	1879
Hydrogen	H	1	0.00009	1766
Indium	In	115	7.31	1863
Iodine	I	127	4.93	1811
Iridium	Ir	192	22.6	1803
Iron	Fe	56	7.87	Ancient
Krypton	Kr	84	0.0037	1898
Lanthanum	La	139	6.15	1839
Lead	Pb	207	11.3	Ancient
Lithium	Li	7	0.53	1817
Lutetium	Lu	175	9.84	1907
Magnesium	Mg	24.5	1.74	1808

\*The density of carbon as graphite is 2.27 g cm<sup>-3</sup>  
The density of carbon as diamond is 3.51 g cm<sup>-3</sup>

Element	Symbol	Relative atomic mass	Density (g cm <sup>-3</sup> )	Date of Discovery
Manganese	Mn	55	7.30	1774
Mercury	Hg	200.5	13.5	Ancient
Molybdenum	Mo	96	10.2	1778
Neodymium	Nd	144	7.01	1885
Neon	Ne	20	0.0009	1898
Neptunium	Np	237	20.2	1940
Nickel	Ni	58.5	8.90	1751
Niobium	Nb	93	8.57	1801
Nitrogen	N	14	0.0013	1772
Osmium	Os	190	22.6	1803
Oxygen	O	16	0.0014	1774
Palladium	Pd	106.5	12.0	1803
Phosphorus	P	31	1.82	1669
Platinum	Pt	195	21.5	1735
Plutonium	Pu	244	19.7	1941
Polonium	Po	209	9.20	1898
Potassium	K	39	0.89	1807
Praseodymium	Pr	141	6.77	1885
Promethium	Pm	145	7.26	1944
Protactinium	Pa	231	15.4	1913
Radium	Ra	226	5.00	1898
Radon	Rn	222	0.0097	1900
Rhenium	Re	186	20.8	1925
Rhodium	Rh	103	12.4	1803
Rubidium	Rb	85.5	1.53	1861
Ruthenium	Ru	101	12.1	1844
Samarium	Sm	150.5	7.52	1853
Scandium	Sc	45	2.99	1879
Selenium	Se	79	4.81	1817
Silicon	Si	28	2.33	1824
Silver	Ag	108	10.5	Ancient
Sodium	Na	23	0.97	1807
Strontium	Sr	87.5	2.64	1790
Sulfur	S	32	2.00	Ancient
Tantalum	Ta	181	16.4	1802
Technetium	Tc	98	11.0	1937
Tellurium	Te	127.5	6.23	1782
Terbium	Tb	159	8.23	1843
Thallium	Tl	204.5	11.8	1861
Thorium	Th	232	11.7	1828
Thulium	Tm	169	9.32	1879
Tin	Sn	118.5	7.29	Ancient
Titanium	Ti	48	4.51	1791
Tungsten	W	184	19.3	1783
Uranium	U	238	19.1	1789
Vanadium	V	51	6.00	1801
Xenon	Xe	131.5	0.0059	1898
Ytterbium	Yb	173	6.90	1878
Yttrium	Y	89	4.47	1789
Zinc	Zn	65.5	7.13	Ancient
Zirconium	Zr	91	6.52	1789

## Formulae of Selected Ions containing more than one kind of Atom

one positive		one negative		two negative		three negative	
Ion	Formula	Ion	Formula	Ion	Formula	Ion	Formula
ammonium	$\text{NH}_4^+$	ethanoate	$\text{CH}_3\text{COO}^-$	carbonate	$\text{CO}_3^{2-}$	phosphate	$\text{PO}_4^{3-}$
		hydrogencarbonate	$\text{HCO}_3^-$	chromate	$\text{CrO}_4^{2-}$		
		hydrogensulfate	$\text{HSO}_4^-$	dichromate	$\text{Cr}_2\text{O}_7^{2-}$		
		hydrogensulfite	$\text{HSO}_3^-$	sulfate	$\text{SO}_4^{2-}$		
		hydroxide	$\text{OH}^-$	sulfite	$\text{SO}_3^{2-}$		
		nitrate	$\text{NO}_3^-$	thiosulfate	$\text{S}_2\text{O}_3^{2-}$		
		permanganate	$\text{MnO}_4^-$				

## Solubilities of Selected Compounds in Water

The table shows how some compounds behave in cold water

vs means very soluble (a solubility greater than  $10 \text{ g l}^{-1}$ )

s means soluble (a solubility of between 1 and  $10 \text{ g l}^{-1}$ )

i means insoluble (a solubility of less than  $1 \text{ g l}^{-1}$ )

— no data

	bromide	carbonate	chloride	iodide	nitrate	phosphate	sulfate	oxide	hydroxide
aluminium	vs	—	vs	vs	vs	i	vs	i	i
ammonium	vs	vs	vs	vs	vs	vs	vs	—	—
barium	vs	i	vs	vs	vs	i	i	vs	vs
calcium	vs	i	vs	vs	vs	i	s	s	s
copper(II)	vs	i	vs	—	vs	i	vs	i	i
iron(II)	vs	i	vs	vs	vs	i	vs	i	i
iron(III)	vs	—	vs	—	vs	i	vs	i	i
lead(II)	s	i	s	i	vs	i	i	i	i
lithium	vs	vs	vs	vs	vs	i	vs	vs	vs
magnesium	vs	i	vs	vs	vs	i	vs	i	i
nickel	vs	i	vs	vs	vs	i	vs	i	i
potassium	vs	vs	vs	vs	vs	vs	vs	vs	vs
silver	i	i	i	i	vs	i	s	i	—
sodium	vs	vs	vs	vs	vs	vs	vs	vs	vs
tin(II)	vs	i	vs	s	—	i	vs	i	i
zinc	vs	i	vs	vs	vs	i	vs	i	i

Note: Some of the compounds in the table hydrolyse significantly in water.



## Melting and Boiling Points of Selected Inorganic Compounds

COVALENT		
Name of compound	mp (°C)	bp (°C)
ammonia	-78	-33
carbon dioxide	-57	-78
carbon monoxide	-205	-192
nitrogen dioxide	-9	21
silicon dioxide	1713	2950
sulfur dioxide	-75	-10
water	0	100

IONIC		
Name of compound	mp (°C)	bp (°C)
barium chloride	961	1560
calcium oxide	2613	2850
lithium bromide	550	1265
magnesium chloride	714	1412
potassium iodide	681	1323
sodium chloride	802	1465

Under normal conditions, carbon dioxide does not melt but sublimates instead. The melting point and boiling point were measured under different conditions.

## Melting and Boiling Points of Selected Organic Compounds

Name of compound	mp (°C)	bp (°C)
methane	-182	-162
ethane	-183	-89
propane	-188	-42
butane	-138	-1
pentane	-130	36
hexane	-95	69
heptane	-91	98
octane	-57	126
cyclobutane	-90	12
cyclopentane	-93	49
cyclohexane	7	81

2-methylpropane	-159	-12
2-methylbutane	-160	28
2-methylpentane	-154	60
2-methylhexane	-118	90

Name of compound	mp (°C)	bp (°C)
ethene	-169	-104
propene	-185	-48
but-1-ene	-185	-6
pent-1-ene	-165	30
hex-1-ene	-140	63

2-methylpropene	-141	-7
2-methylbut-1-ene	-137	31
2-methylpent-1-ene	-136	62
2-methylhex-1-ene	-103	92

methanol	-97.5	65
ethanol	-114	78
propan-1-ol	-124	97
propan-2-ol	-88	82
butan-1-ol	-89	118
butan-2-ol	-89	100

methanoic acid	8	101
ethanoic acid	17	118
propanoic acid	-20	141
butanoic acid	-5	164

### Electrochemical Series (Reduction Reactions)

Metal	Reaction
lithium	$\text{Li}^+(\text{aq}) + \text{e}^- \longrightarrow \text{Li}(\text{s})$
potassium	$\text{K}^+(\text{aq}) + \text{e}^- \longrightarrow \text{K}(\text{s})$
calcium	$\text{Ca}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Ca}(\text{s})$
sodium	$\text{Na}^+(\text{aq}) + \text{e}^- \longrightarrow \text{Na}(\text{s})$
magnesium	$\text{Mg}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Mg}(\text{s})$
aluminium	$\text{Al}^{3+}(\text{aq}) + 3\text{e}^- \longrightarrow \text{Al}(\text{s})$
zinc	$\text{Zn}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Zn}(\text{s})$
iron	$\text{Fe}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Fe}(\text{s})$
nickel	$\text{Ni}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Ni}(\text{s})$
tin	$\text{Sn}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Sn}(\text{s})$
lead	$\text{Pb}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Pb}(\text{s})$
	$\text{Fe}^{3+}(\text{aq}) + 3\text{e}^- \longrightarrow \text{Fe}(\text{s})$
<b>hydrogen</b>	$2\text{H}^+(\text{aq}) + 2\text{e}^- \longrightarrow \text{H}_2(\text{g})$
	$\text{S}_4\text{O}_6^{2-}(\text{aq}) + 2\text{e}^- \longrightarrow 2\text{S}_2\text{O}_3^{2-}(\text{aq})$
	$\text{SO}_4^{2-}(\text{aq}) + 2\text{H}^+(\text{aq}) + 2\text{e}^- \longrightarrow \text{SO}_3^{2-}(\text{aq}) + \text{H}_2\text{O}(\ell)$
copper	$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Cu}(\text{s})$
	$2\text{H}_2\text{O}(\ell) + \text{O}_2(\text{g}) + 4\text{e}^- \longrightarrow 4\text{OH}^-(\text{aq})$
	$\text{I}_2(\text{s}) + 2\text{e}^- \longrightarrow 2\text{I}^-(\text{aq})$
	$\text{Fe}^{3+}(\text{aq}) + \text{e}^- \longrightarrow \text{Fe}^{2+}(\text{aq})$
silver	$\text{Ag}^+(\text{aq}) + \text{e}^- \longrightarrow \text{Ag}(\text{s})$
mercury	$\text{Hg}^{2+}(\text{aq}) + 2\text{e}^- \longrightarrow \text{Hg}(\ell)$
	$\text{Br}_2(\ell) + 2\text{e}^- \longrightarrow 2\text{Br}^-(\text{aq})$
	$\text{Cl}_2(\text{g}) + 2\text{e}^- \longrightarrow 2\text{Cl}^-(\text{aq})$
gold	$\text{Au}^+(\text{aq}) + \text{e}^- \longrightarrow \text{Au}(\text{s})$
	$\text{H}_2\text{O}_2(\text{aq}) + 2\text{H}^+(\text{aq}) + 2\text{e}^- \longrightarrow 2\text{H}_2\text{O}(\ell)$

**Published:** June 2021

**Change since last published:**

House style updates.

Data values updated in line with CRC Handbook of Chemistry and Physics 100<sup>th</sup> ed.

Update of elements on Page 04.